

Abstract

In this experimental study, the most important topics were highlighted, namely that in all oil refineries and oil and gas producing fields, large quantities of gas are burned for purposes including getting rid of sulfur compounds and hydrogen sulfide gas. The amount of gas burned annually is 17 billion cubic meters.. That is equivalent to \$10 billion annually, The process of reducing gas and reusing it has an economic return, as well as preserving the environment and reducing emissions.

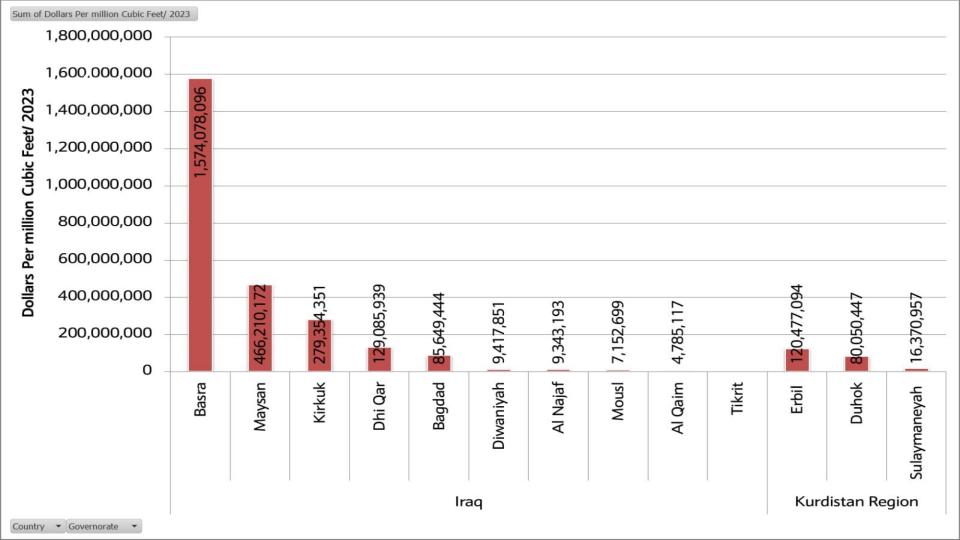




All the metals except gold and platinum combine with sulfur to form inorganic sulfides. Such sulfides are ionic compounds containing the negatively charged sulfide ion S2—; these compounds may be considered as salts of hydrogen sulfide

The sulfur removal rate in this prototype system unit, according to laboratory experiments, is estimated at 98%





Country: Iraq

Flare Type

Upstream

×

Total

3.087627

Number of Flaring Records: 1415 (the same flaring source may have several records)

2013

13.267921

2014

13.983211

2015

16.148394

2016

17.888408

2012

12.641430

Num Sites

1364

Other	51	0.572550	0.424382	0.256040	0.180885	0.184678	0.130190	0.185050
Total	1415	13.213980	13.692303	14.239251	16.329279	18.073086	17.973233	18.006299
Download								
		Flare	Volume in b	illion cubic	meters (BC	(M)		
					Other*			
20			Ops	J Cam	Outer			
18								
16								-
14								-
12		-	-	-				
10								
8								
6								
2								
0								
2012	2013	2014 2	2015 2016	2017	2018 2019	2020	2021 2022	2
	* Othe	r includes refin	eries, downstre	am oil, downstr	ream gas, mids	tream flaring, a	ind LNG.	

17.919503 17.373786 17.800343 17.901724 0.254026 0.227302 0.362495

2020

17.601088

2017

17.843043

2018

17.821249

2019

18.173529

18.162839

2021

180.589012 0.310028

2022

183.676640 18.211752



The system included three towers for stripping, reducing, and adsorbing sulfur from gases using the same catalyst. Depending on the chemical properties of sulfur reactions with metals and the percentage of sulfur in the gas as the chemical reaction sulfur and metals formula

$$F e + S \rightarrow F e S$$
.

$$RS+Fe \rightarrow FeS+R$$
 (R= Mercaptan)



Figure. The flares in the refinery



Figure of prototype system Unit sulfur reducing

Unit Design

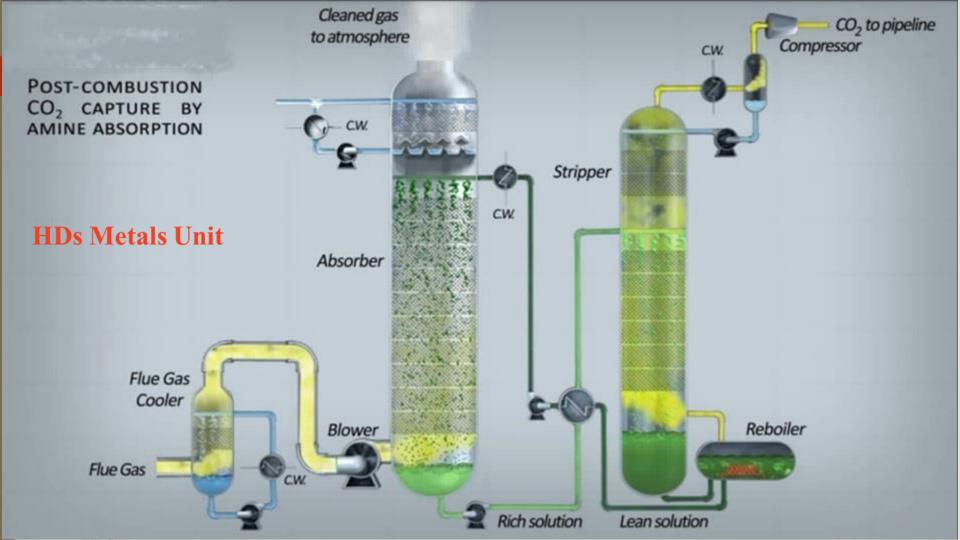
Rotary adsorption tower Paked adsorption tower Reduction tower



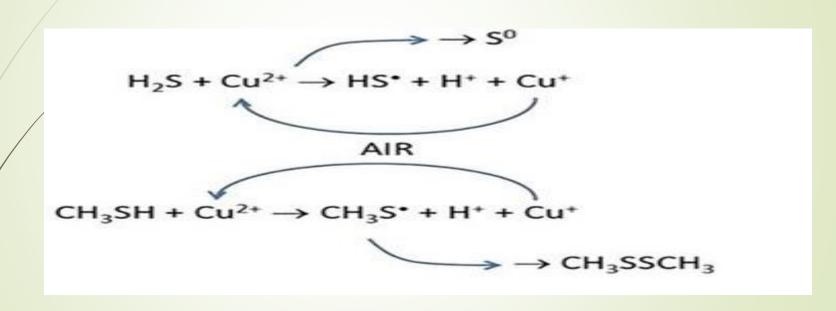
Parts and fillings of towers







Chemical reaction



Thanks For Listening

